$C_2H_4O_3$

7729 – IV

1884

de Forcrand

3. Ann. chim. phys. 3., 187 (1884)

$C_2H_3O_3^-$, s, Hf^o,

$C_2H_4O_3$, s, Hf^o

$C_2H_7O_3N$ (ammonium glycollate), cr, s, Hf^o M, Be

IV 9004

1893

Stohmann. Kleber. Langbein.
and Offenhauer
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$C_2H_4O_3$; rp., ΔH_f°

Circ.500

IV — 9005

1894

Stohmann, Kleber, Langbein,
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J.prakt.Chem.49, 99 (1894)

$C_2H_4O_3$; kp., ΔH_f°

Circ.500 M

$C_2H_4O_3$

30-8482-IV	1906

Muller M. J-A.

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1936

$CH_2OH \cdot COOH$ ($K_{иониз.}$; $\Delta H_{иониз.}$; $\Delta F_{иониз.}$; $\Delta S_{иониз.}$)

Nims L. F.

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B √ ☭

$C_2H_3O_3$

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$C_2H_4O_3$

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Б

Eg2 4.K.

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1953

$HCOOOH$, CH_3COOOH, H_2O_2, CH_3OOH, C_2H_5OOH

(K дисс.)

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В

$C_2H_4O_3$

CH_3COOH (Kp. ΔH) IV – 4280 BP 1960

2 CH_3COOH + H_2O_2 = [redacted] (Kp. ΔH)

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РЖ Хим., 1960, 95580

M, B

3

Ср

CH_3COOOH

$C_2H_4O_3$

(№ 0)

| Bgp-9856-IV | 1963

Schmidt C., Selon A.H.

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CH_3COOOH

ВФ-М1457-IV

1966

ΔH°_{273} =

= −93,330 ккал/моль

ΔU°_{273} = −91,430 ккал/моль

Explosion properties of peroxyacetic acid. I. Thermodynamic calculation of the explosion characteristics of peroxyacetic acid. Stanislav Havel and Jiri Greschner (Vysoka Skola Chem.-Technol., Pardubice, Czech.). *Chem. Prumysl* 16(2), 73–8(1966)(Czech); cf. *CA* 54, 1250*d*. The heat of formation of liquid AcOOH was detd. from the temp. dependence of the AcOH oxidn. equil. const.; ΔH°_{273} = −93,330 kcal./mole and ΔU°_{273} = −91,430 kcal./mole. By using ΔU°_{273} the explosion characteristics were calcd., assuming the formation of CO, CO_2, H_2O, H, and solid C. The calcd. heat of explosion is Q_{273} (H_2O as liquid) = 900 kcal./kg. AcOOH, temp. of explosion t = 1890°, max. thermochem. pressure p = 20,100 atm., vol. of gases V = 898 l./kg., force f = 0.7345 × 10^5 kg./kg., potential A = 29015 kg./kg., and charge d. = 1.1037 g./ml. The detonation wave was calcd. according to the von Neumann-Zeldovitch model (cf. Taylor, *Detonation in Condensed Explosives*, Oxford: Engl., 1952).

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HO CH_3COO

$C_2H_4O_3$

(ΔH_{gaccos})

1971

Avedikian, Leron;

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~~$CH_2OHCOOH$~~

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