

C₂H₄O₃

7729 - IV

1884

de Forcrand

3. Ann. chim. phys. 3., 187 (1884)

$C_2H_3O_3^-$, s., Hf^O ,

$C_2H_4O_3$, s., Hf^O

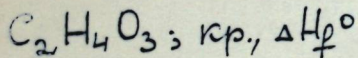
$C_{21}H_{70}O_3N$ (ammonium glycollate),

cr, s, Hf^O M, Be

IV 9004

1893

Stohmann, Kleber, Langbein,
and Offenbauer
2. Ber. Verhandl. sächs. Akad. Wiss.
Leipzig Math. phys. Klasse 45, 604 (1893)



Circ. 500



u.

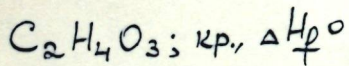
ϕ

IV - 9005

1894

Stohmann, Kleber, Langbein,
and Offenhauer

1.J.prakt.Chem.49, 99 (1894)



Circ.500 M

C₂H₄O₃

BD-8482-IV | 1906

Muller M. F.-A.

Ann. Chim. Phys.

1906, 9, 263-71

CH₂OH·COOH (K_{иониз.}; $\Delta H_{иониз.}$; $\Delta F_{иониз.}$; $\Delta S_{иониз.}$)

IV-4281

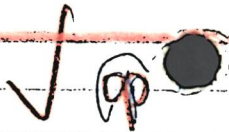
1936

Nims L. F.

J. Am. Chem. Soc. 1936, 58, 987-9

The ionization constant of glycolic acid from 0° to 50°."

B



C₂H₃O₃
C.A., 1936, 58594

943-IV

P, T_m (CH₃COOOH)

1951

Egerton A.C., Emite W., Miukoff G.J.
Discussions Faraday Soc. 1951, No 10,
272-32

"Some properties of organic peroxides

C₂H₄O₃

C.A., 1952, 3358 g

Σ 92 + K.

Б

11 820

1953

HCOOOH , CH_3COOOH , H_2O_2 , CH_3OOH , $\text{C}_2\text{H}_5\text{OOH}$

(K_{prec.})

Everett A.J., Minkoff G.J.

Trans. Faraday Soc., 1953, 49, N 394,
part 4, 410-414 (Ann.)

The dissociation constants of some
alkyl and acyl hydroperoxides.

PX., 1954, N 1,
10326

13

$\text{C}_2\text{H}_4\text{O}_3$

CH_3COOH (кр. дн) IV-4280 BP 1960

~~CH_3COOH (кр. дн) IV-4280 BP 1960~~

Havel S., Weigner J. A.

Chem. průmysl., 1960, 10, №6, 293-296 (revisi.)

Stanovení rovnovážné konstanty v systému
peroxyd vodíku - kyselina octová.

Plh. Xuvn., 1960,
95580

M, B

3

✓ ⊗

CH₃COOH

C₂H₄O₃

(no)

| BQP-9856-IV | 1963

Schmidt C., Selvon A.H.

"Canad. J. Chem.", 1963,
41, 47, 1819-1825

CH₃COOOH

Bp-M1457-IV

1966

$$\Delta H_{273}^{\circ} =$$

$$= -93,330 \frac{\text{kcal.}}{\text{mole}}$$

$$\Delta U_{273}^{\circ} = -91,430 \frac{\text{kcal.}}{\text{mole}}$$

Explosion properties of peroxyacetic acid. I. Thermodynamic calculation of the explosion characteristics of peroxyacetic acid. Stanislav Havel and Jiri Greschner (Vysoka Skola Chem.-Technol., Pardubice, Czech.). *Chem. Prumysl* 16(2), 73-8(1966)(Czech); cf. *CA* 54, 1250d. The heat of formation of liquid AcOOH was detd. from the temp. dependence of the AcOH oxidn. equil. const.; $\Delta H_{273}^{\circ} = -93,330$ kcal./mole and $\Delta U_{273}^{\circ} = -91,430$ kcal./mole. By using ΔU_{273}° the explosion characteristics were calcd., assuming the formation of CO, CO₂, H₂O, H, and solid C. The calcd. heat of explosion is Q_{273} (H₂O as liquid) = 900 kcal./kg. AcOOH, temp. of explosion $t = 1890^{\circ}$, max. thermochem. pressure $p = 20,100$ atm., vol. of gases $V = 898$ l./kg., force $f = 0.7345 \times 10^5$ kg./kg., potential $A = 29015$ kg./kg., and charge d. = 1.1037 g./ml. The detonation wave was calcd. according to the von Neumann-Zeldovitch model (cf. Taylor, *Detonation in Condensed Explosives*, Oxford: Engl., 1952).

T. Boublik

C.A. 1966. 64. 10
14017e

HOCH_2COOH Bps-6955-VI 1969

Bakore G. V. et al.

Kp Z. Phys. Chem (~~SSR~~)

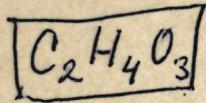
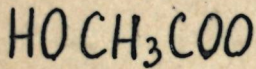
1969, 242, N1-2, 102-6

CH_2OHCOOH Bp-61-XIV 1970

(Kp) Reddy M. V., et al.

J. pract. Chem.

1970, 312, N1, 69-71



(ΔH_{гидролиз})

Arredikiam, Levon;

1971

"J. Chim. Phys. Physicochim
Biol."

1971, 68, (7-8), 1201-5.

● (calc. H₂O; I).

CH_2OHCOOH | Am. 24950 |

1986

~~CH_2OHCOOH~~

u gp.

Hollenstein H., Ha T.-K.,
Bünthard H.H.,

ab initio g. Mol. Struct., 1986, 146,
pacrem 289-307.