

FeOH

Fe-OH

1972

Opitz Ch., Dunker H.H.

Энергия  
и природа  
хим. связи.

"Z. phys. Chem", 1972, 249,  
№ 3-4, 154-160.

(ссыл. Fe-OH, III)

Идентификация  
FeOH

(57-1144)

1973

в минералах  
H<sub>2</sub> + O<sub>2</sub> + N<sub>2</sub>

Jensen David E;  
Jones George A.

соедин.  
железа в  
минералах

"J. Chem. Soc. Far. Trans"  
1973, Part 1, 69, N8,

Fe(OH)<sub>2</sub>,  
FeO



1448-54.

(см. FeOH; I)

1979

$Fe + H_2O$   
 $FeOH$

Hauge R.H. et al.

overcup &  
undercup

NBS Spec. Publ. (U.S.)  
1979, 561-1, 557-66

cup. Mg +  $H_2O$  - III

FeOH

1980

Daidoji H.

J. Spectroscop. Soc. Jap.  
1980, 29, N5, 326-31.

центр  
новиз.

см. Be(OH)<sub>2</sub>-III

FeOH

ommuu 10528

1980

√93: 156599h Thermochemical properties of gaseous iron(II) oxide and iron hydroxide (FeOH). Murad, Edmond (Air

Force Geophys. Lab., Hanscom Air Force Base, MA 01731 USA).  
*J. Chem. Phys.* 1980, 73(3), 1381-5 (Eng). The dissoen. energy of FeOH,  $D_0^0(\text{Fe-OH})$ , was measured mass spectrometrically by passing  $\text{H}_2\text{O}(\text{g})$  and  $\text{H}_2(\text{g})$  over  $\text{Fe}_2\text{O}_3$  in a Knudsen cell. The measurements lead to  $D_0^0(\text{Fe-OH}) = 76.9 \pm 4$  kcal/mol ( $3.3 \pm 0.2$  eV). The ionization potential of FeOH is  $7.9 \pm 0.2$  eV. An equil. involving gaseous FeO was also measured and the results of this study together with an anal. of previous data lead to  $D_0^0(\text{FeO}) = 93.0 \pm 3$  kcal/mol. The following data were derived from the measurement:  $D_0^0(\text{Fe}^+-\text{O}) = 3.15 \pm 0.2$ ,  $D_0^0(\text{Fe}^+-\text{OH}) = 3.3 \pm 0.2$ ,  $D_0^0(\text{FeO}^+-\text{H}) = 4.7 \pm 0.2$ , and  $\text{PA}(\text{FeO}) = D_0^0(\text{FeO}-\text{H}^+) = 9.40 \pm 0.2$  eV. The similarity in bonding between  $D_0^0(\text{Fe-OH})$  and  $D_0^0(\text{Fe-Cl})$  is briefly discussed.

$D_0^0; \gamma$ .

⊠  $\text{Fe}^+-\text{O}$

⊕  $(\gamma^0)$

C.A. 1980, 93, N16

ФедН

130355

1988

Краснов К.С.,  
Филлиппенко М.В.,

ОНИИТЭХИМ.

Дел. N 378-ХП-86,  
Черкасск, 1988.



л.п.

(обзор)

FeOH<sup>+</sup>

(DM 31286)

1988

Tran Q., Kabellas N.S., et al.,

присут-  
ствие в  
пламени,  
масс-  
спектр.

Can. J. Chem. 1988, 66, N 9,  
2216-2218.

Ion Chemistry of transition  
metals in hydrocarbon  
flames. I ● Cation of Fe, Co,  
Ni, Cu and Zn.



FeO<sup>2+</sup>

DM. 37093

1993

McLellough - Catalano J., Lebrilla  
C.B.,

Memor  
quod

J. Amer. Chem. Soc., 1993, 115,  
1447-1448

Determination of  
Bond Energies in



Metal-Hydroxide  
Doubly Charged

$\text{ScOH}$ ,  $\text{FeOH}$  and  $\text{CoOH}$  IONS.



Om. 37095

1993

(OC=1-4)

Schultz R.H., Armentrout P.B.,

J. Phys. Chem., 1993,

97, 596-603

Gas - Phase Metal Ion Ligand -  
Collisions: Induced

Dissociation of  $Fe(H_2O)_x^+$   
and  $Fe(CH_3)_x^+$  ( $x=1-4$ ).