

Li-Fe, Co,
Ni

VI - 2177

1962

$\text{Li}_{0.5}\text{Fe}_{2.5}\text{O}_4$, YbFeO_3 , MgFe_2O_4 ,
 Fe_2O_3 u. sp. (Di)

Wickersheim K., Lefever R.A.,

J. Chem. Phys., 1962, 36, 844-
850

10

1966.

$MgLiBO_3$, $MnLiBO_3$, $CoLiBO_3$,

$ZnLiBO_3$ (Vi)

to $\bar{x}$ 4473

Lehmann H. A., Schadow H.,

Z. Anorg. Allgem. Chem., 1966, 348, (1-2),
42-9.

Chemistry and constitution of borates.

XV. Formation and  preparation

HO

8

γ_1 $[Li_2(UO_2Cl_4), Na_2(UO_2Cl_4),$
 $K_2(UO_2Cl_4), Rb_2(UO_2Cl_4), Cs_2(UO_2Cl_4),$
 $Li_2(UO_2Cl_4) \cdot 2H_2O; Na_2(UO_2Cl_4) \cdot 2H_2O,$
 $K_2(UO_2Cl_4) \cdot 2H_2O; Rb_2(UO_2Cl_4) \cdot 2H_2O; \&$
 $Cs_2(UO_2Cl_4) \cdot 2H_2O)$

Ripani R., Mizel C.

Rev. roumaine chim, 1966, 11, 47, 805-806

Studien über Uranylchloride (neue)

I Thermische Stabilität und IR-Spektren der Uranyl-tetrachlorid-salze mit Alkalimetallen.

Rus. 1964, 12 5188

EC

Ch. H.

VI-7376

1970

Ферроцианиды металлов

Cs, Rb, Ba, Na, Li, Sr, Mg, La, Y, Mg, Zn, Zn
(V;)
и др

Некрасов Б.В., Сейфер Г.В., Харитонов
Изв. АН ССР. Сер. хим., 1970, № 2, 266-
10.9.

271

10

LiCO_2

1970

UK-crekmp

Tarte P., Preud-
homme J.,

Spectrochim. acta,
1970, A26, N4,

747

(all. Li Al
cof)

complexes (Li_3FeF_6 , Na_3FeF_6 , K_3FeF_6 ,
 Rb_3FeF_6 , Cs_3FeF_6 , Ag_3FeF_6 ,
 Tl_3FeF_6) 10 15 16

Shearer - Turrell S., Tressard A.,
X 5301

Pettier J.,
J. mol. Struct., 1971, I, 13-4, 289-300 (ann.)

Structural studies of iron hexafluorides: Infrared spectra of M_3FeF_6
($M = Li, Na, K, Rb, Cs, Ag, Tl$) 14

Pu Xun, 1971, 165217 MO (B)

70127.1217

Ch, TC

40150

Li_2NiO_4

(16f)

1976

5-16513

Sreedharan O.M., Chandrasekharalah
M.S., Karkhanavala M.D. The free energy
of formation of lanthanum nickelate.

"High Temp. Sci.", 1976, 8, N3, 179-183

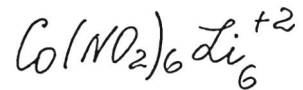
(англ.)

0800 AMK

757 757

292

ВИНИТИ

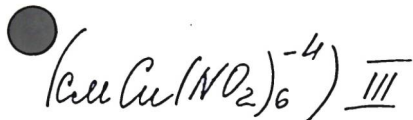


1980

Clark D.W., et al

Kb. Mex.
pacer

Solid State Commun,
1980, 34, 16, 395-99.



Li Fe F₄ ном. 18313 1981

к, инд.
пост. Руднев Е. Б., Борщевский А. Г.,

т. р. Денисов. рук. ВИННТИ,
N 1809-81. Ден., 1981.

Li Fe E₃

10т. 18313/

1981

г, см.
пост.
т. гр. Рудный Е. Б., Борщевский А. Я.,
Депониров. рук. ВИНТИ,
N 1809-81. Дек., 1981.

$\text{LiUF}_6(x)$ [Om. 23747] 1986

Hecht H. G., Malm J. G.
et al.,

(crystal,
vi)

J. Chem. Phys., 1986,
84, N 7, 3653-3662.

LiNiO_2

[Om. 24954]

1986

Itoh M., Yamada T.,
Uwakoshi K., Hirakawa K.,
Yasuoaka H.,

Kp, Tr;

J. Phys. Soc. of Jap.,
1986, ● 55, N 7, 2125 -
2128

$Fe^+ \cdot LiF$

1987

108: 173883s Determination of $R_e(M^{2+}-F^-)$ for chromium(1+), iron(1+), nickel(1+), and manganese(2+) in ionic fluoride lattices from the isotropic superhyperfine constant. Fernandez-Rodrigo, G.; Florez, M.; Pueyo, L.; Moreno, M.; Barriuso, M. *Fac. Quim., Univ. Oviedo, Oviedo, Spain 33007*. *Cryst. Latt. Defects Amorphous Mater.* 1987, 16(1-4), 281-7 (Eng). The det. of equil. metal-ligand distances in transition-metal fluorides from the obsd. isotropic superhyperfine const. $A_{||}$ is analyzed in the light of cluster-type calcs. at the Hartree-Fock-Roothaan level. The $MnFe(4-)$ wavefunctions obtained at different metal-ligand distances confirm that the covalency parameter λ_s is proportional to the $\langle 3d\sigma|2s \rangle$ metal-ligand overlap integral. This behavior is also obsd. in theor. refinements such as the core projection and the cluster-lattice interaction are included into the SCF process. Thus, this result greatly support the method previously used by M. Barriuso and M. Moreno (1984) for deriving the true $Mn^{2+}-F^-$ distances for Mn^{2+} -doped fluoride lattices from obsd. values of $A_{||}$. The present method works also very well for the unstable Cr^+ , Fe^+ , and Ni^+ cations. The $R_e(Fe^+-F^-)$ distances in $Fe^+:LiF$, $Fe^+:NaF$, and $Fe^+:KMgF_3$ are reported here for the first time. Ionic radii and values of $10Dq$ are also given for these unstable systems.

$(r_e(Fe^+-F^-))$

neop. param)

(+2) ☒

C.A. 1988, 108, N20

$Fe^+ \cdot NaF$, $Fe^+ \cdot KMgF_3$

LiUF₆

1989

111: 66732g Optically active vibrations of lithium and α -sodium uranium hexafluorides (LiUF₆ and α -NaUF₆). Ohwada, Ken (Div. Chem., Japan At. Energy Res. Inst., Tokai, Japan). *Appl. Spectrosc.* 1989, 43(4), 714-18 (Eng). Normal coordinate analyses of the optically active vibrations of LiUF₆ and α -NaUF₆ were made, with the assumption of an idealized structural model. On the basis of the results of this treatment, all of the obsd. bands so far reported were successfully assigned to the intramol. vibrations (UF₆⁻) as well as the optically active lattice vibrations [M(Li,Na)...UF₆⁻]. The force consts. concerning U-F and M-F bonds were obtained within the framework of the modified valence force field. The present anal., based on a whole crystal, is powerful not only for assigning the optically active lattice vibrations but also for elucidating the vibronic structure of the electronic absorption spectra obtained at low-temp. measurements.

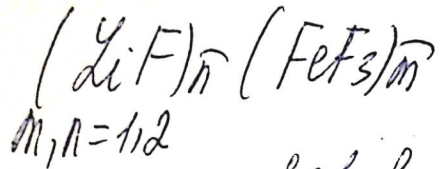
(see norm. Di)

(+2) 

c.A. 1989, 111, N 8



α -NaUF₆
UF₆⁻



1997

Scholz G. et al.,

суп-ра,
стабильн.,
теорет.
расчет

J. Fluorine Chem. 1997,
86(2), 131-42

