

Jr Br x

Отрада и сельскому М. З. | 1956
М. З. З., 1956, 30, №3, 593.

$$-\Delta H_{298}^{\circ}(\text{ZrBz}_2) = 23 \frac{\text{ккал}}{\text{моль}}$$

Теплота
образования
(сущность).

Yr Bz отчет на науче отчета 1965

Ferber R. C.

Rept LA-3164, UC-4
Chemistry. TID-4500
(40th. Ed.)

AMs.

Los Alamos New Mexico, Univ. Ca-
lifor. 1964; distribut may 1965, p. 146

JZ BZ₂

отчет на работе
августов

1965

Ferber. R.C.

Δ Ms. Rept LA-3164, UC-4
Chemistry. TID-4500,
(40th Ed.)

Los Alamos New Mexico, Univ. Cali-
for. 1964; distributed may 1965, p. 145.

YrBz_3

отчет на работе
отчетов

1965

Ferber R.C.

Rept YA-3164, UC-4

ΔMs. Chemistry, TID-4500
(40th Ed.)

Los Alamos New Mexico, Univ. Ca-
lifor. 1964; distributed may 1965, p. 145

$IgBr_3$

1967

1 В52. Новый способ получения безводного трибро-
мида иридия. Колбин Н. И., Самойлов В. М.
«Ж. неорганич. химии», 1967, 12, № 9, 2526
Получен безводный $IgBr_3$ при взаимодействии метал-
лич. Ig с бромом при давлении 8—9 атм при $840^\circ K$.
Автореферат

x. 1968. 2

IrBr₃
p
ΔH_f

101310k Iridium bromides. N. I. Kolbin and V. M. Samoilov (Leningrad. Gos. Univ., Leningrad, USSR). *Zh. Neorg. Khim.* 13(3), 906-8(1968)(Russ). IrBr₃, obtained by bromination of Ir in sealed ampuls at a Br pressure of 8-9 atm. and 840°K., is a cryst., grayish brown powder, with an Ir/Br ratio of 1/3, insol. in H₂O, CCl₄, and C₆H₆. It is stable for 7 days in sealed ampuls at 925°K. and a Br pressure necessary to inhibit the thermal dissoen. Its x-ray pattern shows sharp peaks. The d. is 6.45 g./cc. It is diamagnetic, its sp. magnetic susceptibility K_0 is -0.2×10^{-6} . It begins to dissoc. at 440°C.; the dissoen. pressure at 846°K. is 1 atm.; the heat of formation, $\Delta H_{832^\circ K.}^\circ = 50$ kcal./mole. The product obtained (supposed to be IrBr₃) by Krauss and Gerlach (1925) by treating the blue IrO₂·2H₂O with HBr and heating the reaction product (the supposed IrBr₂) with Br in sealed ampuls on a H₂O bath for 6 hrs. was a dark red-brown powder, x-ray amorphous, paramagnetic, with $K_0 = 0.4 \times 10^{-6}$. Chem. anal. indicates a formula approaching Ir₂OBr₆; the Ir/Br ratio is 1/2.85. The supposed IrBr₂ and the product obtained by its thermal decompn. and designated as IrBr were not examd.

Aniela Klein

1968

6984-11

C. A. 1968. 68. 22

Jr Br 3
cr. str.

V167f0

1968.

Ohnberg H.

Diss. Dokt. Inst. Anorgan. Chem. Friedrich-
Alexander-Univ. Erlangen-Nürnberg,
1968, X 35, 84 S., ill.

Strukturuntersuchungen am Iridium
(III)-bromid und Bestimmung der Fehl-
ordnungserscheinungen bei Edelmetalltri-
halogeniden. Φ U. M. M.
X, 1969, Z 3 B 454.

1969

Вср - 5892 - VI

Уч Вг₃

14 Б630. Термодинамические исследования безводного трибромида иридия. Колбин Н. И., Самойлов В. М. «Ж. неорганической химии», 1969, 14, № 3, 631—635

Исследована термич. диссоциация трибромида иридия в интервале 716—973° К. Измерена теплоемкость $IrBr_3$ при постоянном давлении, $C_p = 99,39 + 20,35 \cdot 10^{-3} T$ дж/моль·град (377—742° К). Определены для р-ции образования $IrBr_3$ из металлич. иридия и газообразного брома: изменение энтальпии $\Delta H^{\circ}_{298,15} = -(223,0 \pm 3,0) \cdot 10^3$ дж/моль, изменение энтропии $\Delta S^{\circ}_{298,15} = -274,0 \pm 4,0$ дж/моль·град и стандартная энтропия тв. $IrBr_3$ $S^{\circ}_{298,15} = 130,5$ дж/моль·град. Резюме

Ср,
8298

к. 1969. 14

IrBr₃

3

C_p ; k given

ΔH_f ; ΔS_f

Bsp - 5892 - VI

1969

109691x Thermodynamic studies of anhydrous iridium tri-bromide. Kolbin, N. I.; Samoilov, V. M. (Leningrad. Gos. Univ., Leningrad, USSR). *Zh. Neorg. Khim.* 1969, 14(3), 631-5 (Russ). The thermal decompn. of IrBr₃ was studied at 716-973°K. in vacuo or in an inert atm. The log of the dissocn. const. was $-10,902/T + 12.88$. At const. pressure, the heat capacity (C_p) of IrBr₃ was $99.39 + 20.35 \times 10^{-3}T$ j./mole-degree at 377-742°K. At 25°, ΔH and ΔS of the formation of IrBr₃ from the elements were $-(223.0 \pm 3) \times 10^3$ j./mole and $-(274.0 \pm 4.0)$ j./mole-degree, resp. The standard entropy of solid IrBr₃ was 130.5 j./mole-degree. HMJR

C.A. 1969. 70. 24

50310.2412

Ph, Ch

~~IrBr₅³⁻~~

IrCl₆³⁻ (англ. текст)

1974

XVI 2258

Sanyal Nitish K., Verma D.N., Dixit L.

Normal coordinate analysis, mean amplitudes of vibration and coriolis coupling

constants of IrCl₆³⁻ and IrBr₅³⁻.

"Indian J. Phys.", 1974, 48, N 6,

481-485

(англ.)

10(Ф)

48, N 6, 41

298

204

0312

ПРИК

ВИНИТИ

To B43

1977

Bain J, et al.

vol. I, p. 333

298-800 (6)

vol. 49; I

IrBr(K)

1984

Pankratz L.B.,

m. sp.
298.15
800K

U.S. Bureau of Mines,
Bull. 674, P. 347.