

Cr-I

CrCl₂

[om. 28273]

1987

Misham M. W. M., Bensen S. W.,

ΔH_f ;

J. Phys. and Chem. Ref.
Data, 1987, 16, N3,

467-470.

$\text{CrICl}_2(\text{H})$ (number 9584) 1974.

$\text{CrIBr}_2(\text{H})$ Cr, Mo, W-compounds.
Barry D.S.

Comput. Anal. Thermo-
chem. data.

(ΔH_f)

CATCh-tables.

Univ. Sussex,

● Brighton.

Sussex, 1974.

CrIO_6^-
 KCrIO_6

[BSP-8299-X]

1973

(Тм, експкт.)
Vi

Kel'iz A., Jüst P.,
C. R. Acad. Sci., 1973,
C 276, NC, 503-506.

Structural study and
thermal decomposition of
tochromate: potassium ioda-

$\text{Cu}(\text{Y}\text{O}_3)_3$

1973

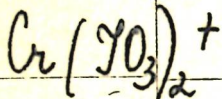
Hassan K., Shleifer J.W., Prescott B.B.

"J. Solid State Chem."

1973. 7, 112, 186-204.

(T₄₂)

■ (see. $\text{Ni}(\text{Y}\text{O}_3)_2 \cdot 4\text{H}_2\text{O}$; I)



BAP-VII-3420

1969

K

~~54274b~~ Complexation of chromium(III) by iodate. Mercer, Edward E.; Hormuth, J. A. (Univ. of South Carolina, Columbia, S.C.). *J. Inorg. Nucl. Chem.* 1969, 31(7), 2145-52 (Eng). The complexes formed between Cr(III) and iodate ion in solns. of ionic strength of 0.5, and $[\text{H}^+] = 0.10M$ were studied. The complexation is very rapid upon mixing of the reagents. Ion-exchange membrane equilibration, spectrophotometric measurements, and soly. studies indicate a single sol. complex, $\text{Cr}(\text{IO}_3)_2^+$, being formed at iodate concns. of $<0.12M$. For the equil. $\text{Cr}^{3+} + 2\text{IO}_3^- \rightleftharpoons \text{Cr}(\text{IO}_3)_2^+$, the best value of K is $132 \pm 10M^{-2}$ at 25° . In addn., the soly. measurements indicated higher complexes are formed at iodate concns. of $>0.2M$.

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C.A. 1969. H. 12