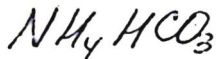


CNH₅O₃

NH₄HCO₃

N1119



T_m

| B.D. - 1840 - IV | 1929

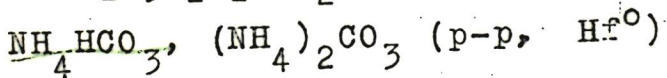
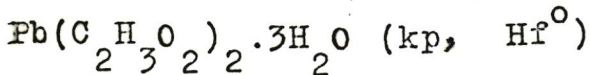
Ferres E.,
Behrens H.

Z. phys. Chem., Abt. A
Haber - Band, 1929,
695 - 716

7268

-117

1873



Berthelot

4 .Ann.chim.phys., 29, 433 (1873)

M, W



1887

7017 - II

NH_4HCO_3 (ag. H)

Berthelot M., Andre

1. Amer. Chim. phys., 1887, 11, 332-41

M, W

CNH_5O_3

7307-10

1899

Berthelot and Andrieu
4. Ann. chim. phys. 17, 433 (1899)

NH_4HCO_3 ; kp. , p-p. , ΔH_f°

$\text{C}_2\text{H}_3\text{ON}$ (mc. , ΔH_f°)

Circ. 500

M 103

CNH_5O_3

~~JP~~

7012 - 16

1926

NH_4HCO_3 , kp, (p)

Bonnier

1. Ann. chim. 5, 37 (1926)

M, W

CNH_5O_3

NH_4HCO_3 (T_m)

1840-IV

1928

Terres G., Behrens H.

Z. physik. Chem., Abt. A, Haber Bd.
139, 695-716 (1928)

"The physical-chemical principles
of urea synthesis from ammonia,
carbon dioxide
and water." B. 51(9) | C.A., 1930, 4691 L

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1929



(cr., Hf)

Kelley

4. J. Amer. Chem. Soc., 51, 1400
1929

M, W





Hutchison Wm. K.

J. Chem. Soc. 1931, 410-2

"Equilibrium constants for the
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bicarbonate."

M

V/q

C.A., 1931, 2629

1841-12

1954

NH_4HCO_3 (p. Tm)

Zernike J.

Rec. trav. chim., 1954, 73, 95. 101

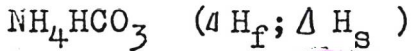
Three-phase curve of ammonium
bicarbonate

Б

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C.A., 1954, 7412 d

8295-17

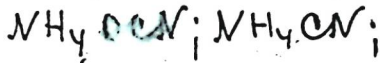
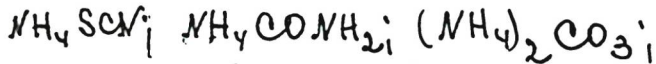


1955



Luft N.W.

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 heats of ammonium salts.

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 75524

 2
 number

БМ

g

11-1838 I958

H_3O^+ , $COONH_4OH$, $COONH_4OH_2^+$ (KP)
 $COONH_4NH_2$, $COONH_4NH_3$ (Kp)

Abu E.M.S., Abdel Halim, Nussein Sadek
Z.phys.Chem.(BRD), 1958, I4, N 3-4, I56-
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The effect of mannitol or urea on dilute
aqueous HCl solutions.

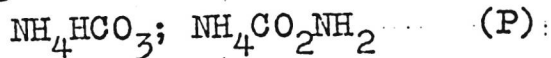
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CNH_5O_3

11

9535 - 12

1961



Barański A.

Chem. Stosowana, 1961, 5, 443-5

The dissociation vapor pressure ..

Be

CN O₂ H₅



1963

 NH_4HCO_3

Thermogravimetric investigation of decomposition of ammonium bicarbonate. Walter Wanek and Jaroslav Nyvlt (Res. Inst. Inorg. Chemistry, Usti n.L., Czech.). *Chem. Prumysl* 13(3), 123-6(1963). Decompn. of NH_4HCO_3 was studied at 20 to 170° and under isothermal conditions at 20° for 25 hrs., at 40° for 25 hrs., and at 60° for 5 hrs. Rate of decompn. was directly proportional to time and to surface of crystals, and increased with increasing temp. and moisture. Activation energy at 70 to 120° was 10.6 kcal./mol. If the reaction is considered to be an evapn., this value should be 2/3 of the heat of evapn., that is 10.2 kcal./mol. Vapor pressure of the salt contg. 0.5-5.0% moisture at 20° was 7.5 times that of the dry substance; 2.5 times owing to water, 5 times owing to the substance proper. Production of large crystals to improve storage properties in agricultural applications is suggested.

B. G. Mandelik

N1119
 CNH_5O_3

C.A. 1963.59.5

4765c

$(\text{NH}_4)_2\text{SO}_3$ (vr. sh)

50304.9042

48536

1974

Ex-Ch/XHB-z, Ch-z, TC

NH_4HCO_3

4036

Czarnocki J., Habor J., Pawlikowska-Czubak J., Pomianowski A. Studies on kinetics and mechanism of the reaction of sulphur dioxide and nitrogen oxides with gaseous ammonia. II. Thermodynamic equilibria in the system SO_2 , NH_3 , H_2O and CO_2 ; Z. anorg. allg. Chem.

287 290

1974, 470, 213-18

05208 mm

ВИНИТИ

NH_4HCO_3

XIII-5668
House J. E. *

1980

ΔH_{tr}

Inorg. and Nucl. Chem.
Lett., 1980, 16, n 4, 185-7



$\text{Cu}(\text{NH}_4)_2\text{CO}_3$ i $\bar{I}$