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85: 83787r Stability of nickel(II) monoammoniate in aqueous solutions of methyl, ethyl, and isopropyl alcohols. Krestov, G. A.; Shormanov, V. A.; Afanas'ev, V. N. (Ivanov. Khim.-Tekhnol. Inst., Ivanovo, USSR). *Zh. Neorg. Khim.* 1976, 21(3), 738-41 (Russ). Stability consts. of Ni^{2+} monoammine complexes were detd. potentiometrically in aq. MeOH, EtOH, or iso-PrOH at 15–35° and ionic strength 1. Thermodn. parameters (ΔH and ΔS) for complexation are small and nearly equal expt. error. Complex stability increases as alc. content increases; plots of log β_1 vs. reciprocal dielec. const. are linear or nearly linear.

V. Justova

C.A. 1976 85 N 12

Комплексы Ni
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1983

Chaudhuri A.K.,
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Ni - сплавы (Cm. 24371) 1986

Brandt R., Pawlowski Z., Neuer G., et al.,

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$N_2 \cdot Ar$ f-p

1988

(G)

110: 161244n Dilute solid solutions of nitrogen-14, nitrogen-15, and carbon monoxide in argon at helium temperatures: heat capacities and rotational spectra of impurity molecules. Krivchikov, A. I.; Bagatskii, M. I.; Manzhelii, V. G.; Minchina, I. Ya.; Muromtsev, P. I. (Fiz.-Tekh. Inst. Nizk. Temp., Kharkov, USSR). *Fiz. Nizk. Temp. (Kiev)* 1988, 14(11), 1208–12 (Russ). The heat capacities were detd. of dil. solid solns. of $^{15}N_2$ and CO in Ar at 0.5-6 K, for impurity concns. of ≤0.25 mol.%. Energy spectra of retarded rotational motion of N_2 and CO mols. in solid Ar are defined.

(H) ⊠ $CO \cdot Ar$

C.A. 1989, 110, N 18

Сплавы Ni

(Δ, K.)

1988

111: 46027c Enthalpies of formation of liquid and solid binary alloys based on 3d metals. V. Alloys of nickel. Niessen, A. K.; Miedema, A. R.; De Boer, F. R.; Boom, R. (Philips Res. Lab., 5600 NA Eindhoven, Neth.). *Physica B (Amsterdam)* 1988, 152(3), 303–46 (Eng). In continuation of the papers on alloys based on either Sc, Ti, V, Cr, Mn, Fe or Co, a review was made, on the basis of the model developed by Miedema and co-workers, of enthalpies of formation of ordered binary intermetallic compds. of Ni with arbitrary metal partners. The enthalpy effects are also examd. for liq. Ni alloys. The calcd. values agree quite satisfactorily with the available exptl. data for binary systems of Ni with a transition metal, except where Ni is the minority partner in alloys with the much more electropos. metals like Y, La and Ti, in which case the model tends to overestimate the enthalpy effects. A formalism is suggested to improve the predictions in these cases. In the binary systems where Ni is alloyed with a nontransition metal, differences are sometimes obsd. between predicted and exptl. detd. enthalpies. These discrepancies are analyzed by taking into consideration information on all 3d metals. The limitations of the model will become clear for Ni alloys in particular.

(ΔfH)

C.A. 1989, 111, N 6

Ni - комплекси

1989

114: 109644m Equilibria between low-spin state (D_{4h}) and high-spin state (O_h) of macrocyclic ligand complex ion (NiL^{2+}). Byun, Jong Chul; Han, Sung Sin (Nat. Sci. Coll., Cheju Univ., Cheju, S. Korea). *Nonmunjip - Cheju Taehakkyo, Chayon Kwahakpyon* 1989, 29, 191-9 (Korean). The chem. equil. of Ni(II)-tetraamine(tetraamine = L, 2, 12-dimethyl-3.7.11.17-tetraazabicyclo-11,3,1-heptadeca-1(17), 2,11,13,15-pentaene) complex ion with water, acetonitrile, acetone and nitromethane were investigated by using a spectrophotometric method. The equil. between low-spin(D_{4h}) and high-spin(O_h) structures of Ni-tetraamine complex ion are presented in water, acetonitrile and acetone, but not in nitromethane. The equil. consts., the reaction enthalpies and the reaction entropies were detd. from anal. of the temp. dependence of the electronic spectra. The formation of the triplet species (O_h) was found to be exothermic. The solvent and electrolyte effects on the equil. consts. could be explained by the dielec. consts. of solvents and the reaction entropies.

(Kp)

C.A. 1991, 114, N12.

Слава Ni 1989

Цанков В. И., Андреев
В. О. и др.

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