

P₂



P₂

Ганс Э

Кубачевский О.

1959

М-х в металлургии

$S_{298} = 52,15 \pm 0,2$ (316) Данные
взяты у Zeise & Electroch.
1542, 48, 476

1959

1959

1959

P₂:

из-за износа
до 20000 K

ВР - 10012 - III

1936

Горел

Верхний

ис. 9 X

8, 904

Ср, S и 9

использ P₂

P₂

1941

ing given

Stevenson D.P., *Proc Roy Soc A*
J. Chem. Phys. 9, 403

Термодинам. р-ции. распада, расщепления
 и некоторых реакций распада

расщепл. $\omega_{11} = 770$ $\delta = 2$: $P-P = 1,88 \text{ \AA}$
 $J = 50,47 \cdot 10^{-80}$, $X' \Sigma$, $\omega = (n+1/2) 780,76 -$
 $-(n+1/2)^2 2,98$ (●) Φ_{11} ω_{11} в cm^{-1}

vs International Critical Tables

29816	9px	5
400	45,00	52,15
600	47,13	-
800	50,17	-
1000	52,39	-
1200	54,18	-
1400	55,65	-
1500	56,92	-
1500	57,50	-

4

I950

P₂

Prever L.

S_T^O

298-2000°K

The chemistry and metallurgy of miscellaneous materials. Thermodynamics

I950,p. I3

19.50

P₂(2)

Kelley KK

Bull WFF

S₂₉₈

$$S_{298} = 52,11 \pm 0,10$$

P₂

Stull D.R., Sinke G.C. 1958

m-g The Thermodynamic Properties
of the Elements
1958

Вр - 1364 - III

1961

P₂

Potter R. Z., Di Stefano V. N.

н-з глицерин
до 5000% Термодинамические функции
некоторых фосфорных соединений. —
J. Phys. Chem., 1961, 65, 849.

см. P

P₂

гас

т.ф.

1962

губерн Л. В. и др.

Москва, 1962

Кертодидиасириесии
св-ва индивидуаль-
ных версий.

P2(203)

McBride B. & G. P.

1963

Thermodynamic properties...
NASA SP-3001, Washington, 1963

$$M = 61,950; \sigma = 2$$

To we were we be α_1 α_2 De 1

$M'Z$ 0 780,89 2,820-0,0051/0,30359 0,444 $\cdot 10^{-2}$ $2 \cdot 10^{-6}$ [0,184] 10^{-6}
A' Π 344343 618,88 [2,97] - 0,27520 $1,69 \cdot 10^{-3}$ [0,218] 10^{-6} 2

	CP
298,15	7,6561
3000	9,0839
6000	9,3678

H-H
2128,2
26027,5
53633,6

S
52,1076
72,0427
78,4106

P_2 (ray)

7ANAT

1965

T. ϕ .

100 - 6000°K



1968

P_2 P_4

niepnoz.
 ps - yue

-95687g Thermodynamic properties of P_2 , P_4 , and some phosphorus fluorides. O'Hare, P. A. G. (Argonne Nat. Lab., Argonne, Ill.). U.S. At. Energy Comm. 1968, ANL-7459, 29 pp. (Eng). Avail. Dep.; CFSTI. From *Nucl. Sci. Abstr.* 1969, 23(12), 24056. The thermodynamic functions S^0 , C_p^0 , $(H^0 - H_0^0)/T$, $(G^0 - H_0^0)/T$, and $(H^0 - H_{298}^0)$ were calcd. for the ideal gas mols. P_2 , P_4 , PF , PF_2 , PF_3 , PF_4 , PF_5 , and P_2F_4 , by the formulas of statistical mechanics. The quantities ΔH_f , ΔG_f , and $\log K_f$, are also given; where exptl. values were not available, est. were made. The thermodynamic data were calcd. at selected intervals 0-6000°K., and also at 273.15 and 298.15°K. Some thermodynamic aspects of the reactions of P fluorides are discussed.

TCNG

C.A. 1969.

71.20

77



P_2
(Ideal gas)

100 - 6000°K
(196)

FINAL
Temp

197

ammek 2502

1968

P₂, P₄,

P. A. F. O'Hare

m.g. op.

Argonne National Labor.

9700 South Cass Avenue

ANL - 7459, Chem.

Dec. 1968, 1-29.

P_2^+

ammuck 3176

1974

1000 - 9400°K

Schneider J

1000 - 8200°K

Z. Phys. Chem

1974 255, N5, 986-96

P₂

ammun 3176

1974

Schneider J.

1000 - 13100°K

1000 - 8400°K

Z. Phys. Chem

1974, 255, N5, 986-96

P₂

1977

Barin J, et al.

~~v. I; p. 591~~

v. II; p. 505

298-2000₂.

(cur. Ag-I)

P₂ (2) Турбин Л.В. и др. 1978

Термогидроэнергетика
м. ф. инж. в-в, 3-е изд. м.т.
стр. 264.

И., Москва, 1978

P₂(2)

1982

Par Kratz L.B.

(198-2000) Thermodynamic Properties
of Elements and Oxides
USA Bur. Mines Bull. 672

● (yul'eghecha)