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Recent investigations have suggested that the long-held belief that Ca oxide has a more neg. Gibbs energy of formation than Mg oxide at lower temps. is inaccurate. If these claims are valid, elemental Mg should be an effective reducing agent for CaO. To test this hypothesis, mixts. of elemental Mg and CaO and of elemental Ca and MgO were sealed under vacuum in molybdenum crucibles and equilibrated at 1398-1473 K for extended periods of time. Although anal. results suggest that true equil. was not established in these expts., the data do suggest that the needed correction of the "traditional" thermodn. data for CaO and MgO is not as great as claimed.