

C4 N3

$C(CN)_3^-$

Baer W. K.

1962

Chem

Diss. Abstr., 1962, 25, n6, 1987.

ИК - спектры и спектры
комб. масс. $VOBr_3$, $AsBr_3$,
 $PO(NCO)_3$, $P(NCO)_3$, $ClC(CN)_3$
 $[C(CN)_3]^-$ - иона.

(сое. $VOBr_3$)

C. A. 1963. 58 n8

7515f

1962

 $C(CN)_3^-$

6 Д223. Колебательные спектры и структура иона трицианометанида $C(CN)_3^-$. Long D. A., Carrington R. A. G., Gravenor R. B. Vibrational spectra and structure of the tricyanomethanide ion $C(CN)_3^-$. «Nature» (Engl.), 1962, 196, № 4852, 371—372 (англ.)

На основании изучения ИК-спектра и спектра комб. рас. натриевой соли трицианометанида сделан вывод о плоской структуре иона.

ф. 1963. 69

$[C(CN)_3]^-$

Miller F. A.,
Baer W. K.

1963

спектр.

Spectrochim. acta, 1963,
19, n 1, 73.

$C_2N_3^-$

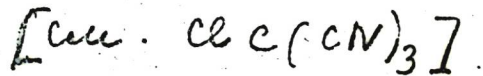
ИК-спектры и спектры
комбинационного рассея-
ния



хлорэтириума
металла на и иона

ф. 1966. 67

трицианметана.



C(CN)₃

Bp-11518-IV 1963

Miller F. A

(vi)

Pure and Appl.

Chem., 1963, 4 n1, 125-9

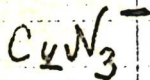
$C(CN)_3$

(vi)

1963



$$V_i$$

$$u.u. \text{ u.u.}$$


Potential constants of tricyanomethanide ion. G. Nagarajan (Oklahoma State Univ., Stillwater). *Indian J. Pure Appl. Phys.* 1(7), 273-4(1963); cf. Wilson, *CA* 34, 295^s; 35^o 1280^s; Pistorius, *CA* 53, 8807g. Nine force consts. for $[C(CN)_3]^-$, possessing a planar sym. structure with the point group D_{3h} , were calcd. on the basis of Wilson's group theoretical method. Their values in millidynes/A. were: f_d 4.658, f_{dd} 1.123, f_r 18.226, f_{rr} 1.443, f_{dr} 1.885, f_α 0.677, f_β 0.493, f_γ 0.068, and f_a 0.149, where f_d is the force const. owing to C—C stretching, f_r that of C:N stretching, f_α that of C—C—C bending, f_β that of C—C:N bending, f_γ that of the C—C bond with the plane of the ion, f_a that of the out-of-plane bending between C—C and C:N bonds, and others are the resp. interaction consts. BGJN

$$C.A. 1963.59.12$$

$$13458fg$$

C(CN)₃⁻

Radou L.

1976

Aust. J. Chem. 1976
29(8), 1635-40

(radou
cupyru)

[all NH₂⁻] III

$C(CN)_3^-$ [номер 8581] 1979.

Alix A. Z. P.
et al.

сил. пог.
сп. анал.
колебен.

Spectrosc. Lett.
1979, 12 (5), 387-404





1989

110: 237769k Chemical modeling of the solvation of dicyanamide and tricyanomethanide ions. Garbuz, S. V.; Skopenko, V. V.; Khavryuchenko, B. D.; Gerasimchuk, N. N. (Kiev. Gos. Univ., Kiev, USSR). *Teor. Eksp. Khim.* 1989, 25(1), 92-6 (Russ). The SCF MO-LCAO method and the semi-empirical MNDO-HB (H bond) approxn. were used to calc. geometric and electronic structures of $C(CN)_3^-$ and $N(CN)_2^-$ solvates in alc., H_2O , $CHCl_3$, or CH_2Cl_2 solns. The monosolvates of these ions resemble H bonded complexes of av. stability.

(meop. para)

C.A. 1989, 110, N 26

$C(CN)_3^-$

1995

123: 123711y Molecular constants of tricyanomethanide ion. Kamaraj, V. (Department of Physics, Anna University, Madras, 600 025 India). *Asian J. Chem.* 1995, 7(3), 604-7 (Eng). Coriolis coupling coeffs., centrifugal distortion consts. and thermodyn. functions of tricyanomethanide ion $C(CN)_3^-$ have been calcd. using the force consts. evaluated on the basis of General Valence Force Field.

M.A.,

Chem. room,
paceram

C.A. 1995, 123, N10