Hg – J

V 1631

1941

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HgClBr, HgBrI, HgClI (ω)

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HgI_3^-

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cell. $HgCl_2$ - I

HgClI

HgBrI

HgBrCl

(cneurp)

1980

92: 155279w The electronic spectra of the mixed mercury dihalides. Part 1. Computational procedures for calculating spectra, for a new route to equilibrium and formation constants, and the resolved spectra. Griffiths, Trevor R.; Anderson, Richard A. (Dep. Inorg. Struct. Chem., Univ. Leeds, Leeds, Engl. LS2 9JT). *J. Chem. Soc., Dalton Trans.* 1980, (2), 205-8 (Eng). Three methods are described for computing the electronic spectra of HgClI, HgBrI, and HgBrCl, free from the contributions of other species, based on the exptl. spectra of equimolar mixts. of HgX_2 and HgY_2 (X, Y = halogen), or of HgX_2 with added Y^-. A procedure is given for computing the formation consts. of the mixed species; the spectra were resolved into their component bands, and their spectral parameters were calcd. Each spectrum contains 3 bands, with peak max. at 37,800, 46,000, and 53,400 for HgClI, 37,800, 43,200, and 51,210 for HgBrI, and 43,500, 47,400, and 57,200 cm^{-1} for HgBrCl.

C.A. 1980. 92 N18

HgClI

HgClBr

HgBrI

1980

92: 155280q The electronic spectra of the mixed mercury dihalides. Part 2. Identification, equilibrium and formation constants, and assignment of transitions. Griffiths, Trevor R.; Anderson, Richard A. (Dep. Inorg. Struct. Chem., Univ. Leeds, Leeds, Engl. LS2 9JT). *J. Chem. Soc., Dalton Trans.* 1980, (2), 209–15 (Eng). The electronic absorption spectra of HgClI, HgBrI, and HgBrCl in H_2O at 20° were computed at 200–320 nm; they are not the mean of the spectra of HgX_2 + HgY_2 (X, Y = halogen). For the reaction $HgX_2 + HgY_2 \rightleftarrows 2HgXY$, the spectra were derived by 2 different methods which gave the same profile. Equil. consts. (log *K*) for this exothermic reaction, independent of added $NaClO_4$ are: HgClI 1.40, HgBrI 1.26, and HgBrCl 0.70. Values of the equil. const. log *K* for $HgX_2 + Y^- \rightleftarrows HgXY + Y^-$, obtained by a graphical method, are 0.98, 0.52, and 0.23 for XY = ClI, BrI, and BrCl, resp. The spectra calcd. by this method agree with those calcd. by the other 2 methods. The resolved spectral bands were assigned and explained.

Del. Chem. J.

C.A. 1980. 92, n 18

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HgJ_3^-

(конд.)

93: 227825s Compounds of complex halo and pseudohalo acids of the Group II B metals. Part III. Spectral studies of the triiodomercurate(II) (HgI_3^-) ion. Perlepes, S. P.; Zafiropoulos, T. F.; Kouinis, J. K.; Galinos, A. G. (Dep. Inorg. Chem., Univ. Patras, Patras, Greece). *Z. Naturforsch., B: Anorg. Chem., Org. Chem.* 1980, 35B(10), 1244-6 (Eng). A study has been made of the far-IR spectra (250-30 cm^{-1}) of the complexes in the solid state. Assocd. anionic structures are present in the 2 complexes. The UV spectrum of the aniline complex, in 3 solvents, show that the HgI_3^- ion is the predominant anionic species in soln.

Эі, ИК спектр

С.А. 1980, 93, N 24

$HgI_4Cl_2^{2-}$

ИК спектр

Оттиск 13126 | 1981

Contreras J. B., Seguel G. V.,

Spectrochim. acta, 1981, A37, N11, 1011–1014.

$HgJBr_2^{2-}$ | Отпеск 18126 | 1981

ИК спектр

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$J_2(B^3\Pi) + Hg \rightarrow JHg + J$ 1983

24 Б1106. Поиск лазерно-индуцированной реакции возбужденного $J_2(B^3\Pi)$ с Hg в скрещенных пучках. Search for the laser-induced crossed beam reaction of excited $J_2(B^3\Pi)$ with Hg. Oprysko M. M., Aoiz F. J., Bernstein R. B., McMahan M. A. «Chem. Phys.», 1983, 79, № 3, 341—350 (англ.)

Сообщается о безуспешной попытке детектирования продукта JHg, возникающего в р-ции $J_2(B^3\Pi) + Hg \rightarrow JHg + J$ в условиях молек. пучков при относит. энергии реагентов $\leqslant 2{,}8$ эВ. Прямое обнаружение JHg в этих условиях затрудняется малой долей возбужденного йода (возбуждение аргоновым лазером) (0,38%), малым телесным углом рассеяния продуктов и малой энергии связи продукта. Сечение р-ции, если она вообще идет, не превышает половины сечения р-ции с участием невозбужденного йода. Поскольку в экспериментах в газе обнаружена р-ция электронно-возбужденного J_2 с Hg, она, вероятно, должна быть интерпретирована как результат многократных столкновений, к-рых нет в экспериментах с пучками.

Е. Е. Никитин

Х. 1983, 19, N 24

$Hg_n J^-(H_2O)_6$

$n = 6, 7$

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Electrochim. acta,
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(сер. $Mg_n Cl^-(H_2O)_6$; III)

HgClJ

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Kaupp Martin, von Schnering Hans Georg.

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М.П.

Р.Ж.Х. N18, 1996, 18 5124